

Organic Chemistry Experiments

Preparation of Manganese(III) Acetyllacetonate





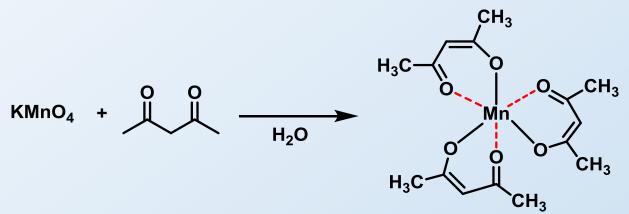
Complex or **Coordination compound**, a class of substances in which a central metal atom is surrounded by other nonmetal atoms. Through the preparation of manganese(III) acetyllacetonate, you can be familiar with the following knowledge:

- 1. Master the synthetic methods of complexes
- 2. Master the operation of mixed solvent recrystallization

3. Review the manipulations of reflux and simple distillation reaction device







 $KMnO_4$ undergoes a redox reaction with acetylacetone, Mn^{VII} is reduced to Mn^{III} , and then coordinated with acetylacetone to afford manganese(III) acetyllacetonate.



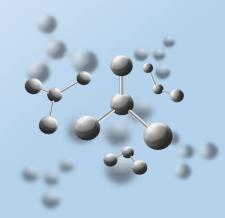
Reagents

Acetylacetone:	CH ₃ COCH ₂ COCH ₃	F.W.	100.11
	b.p. 139 °C	d.	0.976

- Potassium permanganate: $KMnO_4$ F.W. 159.00
- Acetone: CH_3COCH_3 F.W.58.08b.p. $56 \,^{\circ}C$ d.0.788

Petroleum ether: b.p. 60-90 °C

Manganese(III) acetyllacetonate: Mn(C₅H₇O₂)₃ b.p. 60-90 °C



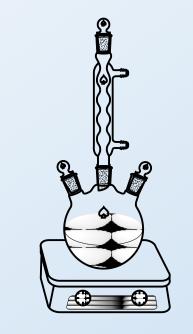


Procedures

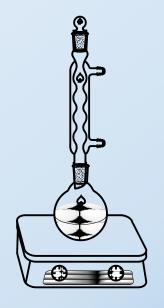
- **1)** KMnO₄ (1 g) + H₂O (10 mL), stirred at 75 °C;
- 2) Add acetylacetone 4.4 g (~ 4.8 mL) dropwise and stir at 75
 °C for 10 min;
- 3) Remove heating, cooled with ice-water bath for 10 min;
- 4) Filtered, washed with H_2O and 30% EtOH, then dried to constant weight;
- 5) Recrystallization with acetone-petroleum ether •



Procedures



♦ Apparatus for preparation







Procedures

◆ preparation



♦ Filteredand washed



Recrystallization



Driedunder IR





Notes

1) The reaction is very violent when initiating. Therefore, acetylacetone needs to be added dropwise.

2) The product has good solubility in ethanol. So the the concentration of EtOH should not be too high during washing.

